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CHEMISTRY

Time allowed

25 Minutes

2002

Topic Questions

3.2 Inorganic chemistry

AQA

AS & A LEVEL

Percentage

%

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Score

/21



Which one of the following electronic configurations is that of a transition element?

- **A** [Ar] 4s²3d¹⁰
- **B** [Ar] 4s²3d⁹
- **C** [A] 4s²3d⁰
- **D** [Ar] $4s^23d^{10}4p^1$





The oxidation of ethanedioate (*oxalate*) ions by manganate(VII) ions can be represented by the half equations:

 $C_2O_4^{2-}(aq) \rightarrow 2CO_2(g) + 2e^{-}$

 $MnO\overline{4}$ (aq) + 8H⁺(aq) + 5e⁻ \rightarrow $Mn^{2+}(aq)$ + 4H₂O(I)

What volume (in cm³) of 0.02 M KMnO₄ is required to oxidise completely a solution containing 0.02 mol of ethanedioate ions?

- **A** 25
- **B** 40
- **C** 250
- **D** 400

(Total 1 mark)



When vanadium reacts with chlorine at 400°C, a brown compound is obtained. When an aqueous solution containing 0.193 g of this compound was treated with aqueous silver nitrate all the chlorine in the compound was precipitated as silver chloride. The mass of silver chloride (AgCl) produced was 0.574 g. Which one of the following could be the formula of the brown compound?

- A VCI
- B VCl₂
- C VCl₃
- D VCl₄



In which one of the following reactions does the metal species undergo reduction?

- $\textbf{A} \qquad MnO_2 + 4HCl \rightarrow MnCl_2 + 2H_2O + Cl_2$
- $\mathbf{B} \qquad [\mathsf{Cu}(\mathsf{H}_2\mathsf{O})_6]^{2+} + 4\mathsf{Cl}^- \rightarrow [\mathsf{Cu}\mathsf{Cl}_4]^{2-} + 6\mathsf{H}_2\mathsf{O}$
- **C** $\operatorname{CrO}_{7}^{2-} + 2\operatorname{OH}^{-} \rightarrow 2\operatorname{CrO}_{4}^{2-} + \operatorname{H}_{2}\operatorname{O}$
- **D** $TiO_2 + 2C + 2CI_2 \rightarrow TiCI_4 + 2CO$

(Total 1 mark)

Which one of the following statements is true?

- **A** A blue solution containing the ion $[CoCl_4]^{2-}$ turns pink when added to an excess of water.
- **B** A purple solution is formed when chlorine is bubbled into aqueous sodium bromide.
- **C** A yellow precipitate is formed when aqueous silver nitrate is added to aqueous sodium chloride.
- **D** A green solution containing the ion $[CuCl_4]^{2-}$ turns blue when added to an excess of concentrated hydrochloric acid.

(Total 1 mark)

6

Which of the species given below can behave as ligands?



- A all three
- B only NH₃
- C NH₃ and NH₄
- D NH₂ and NH₃





The percentage of iron in a sample of impure iron(II) sulphate crystals can be determined by titrating solutions, made from separate weighed samples acidified with dilute sulphuric acid, against a standard solution of potassium manganate(VII).

Which one of the following would lead to the greatest error in the calculation of the percentage of iron(II) in the sample?

- A an error of 0.005 g made when weighing out a sample of mass 0.987 g
- **B** an end-point error of 0.1 cm³ in 25.0 cm³
- **C** an error of 5 cm³ when measuring out 25.0 cm³ of dilute sulphuric acid
- **D** using the average of the titration values 25.4, 25.7 and 25.9 when the correct value is 25.5 cm³

(Total 1 mark)



The percentage of iron in a sample of impure iron(II) sulphate crystals can be determined by titrating solutions, made from separate weighed samples acidified with dilute sulphuric acid, against a standard solution of potassium manganate(VII).

Which one of the following would lead to an inaccurate result?

- A transferring the weighed sample of iron(II) sulphate into a wet conical flask
- **B** failing to measure accurately the volume of water used to dissolve each weighed sample of iron(II) sulphate
- **C** transferring the standard solution of potassium manganate(VII) from its original container to the burette using a wet beaker
- **D** failing to measure accurately the volume of dilute sulphuric acid added to the mixture before titration





The percentage of iron in a sample of impure iron(II) sulphate crystals can be determined by titrating solutions, made from separate weighed samples acidified with dilute sulphuric acid, against a standard solution of potassium manganate(VII).

Which one of the following statements explains why dilute hydrochloric acid is unsuitable for use in this titration?

- A HCl will oxidise Fe²⁺ to Fe³⁺
- B Cl⁻ will reduce Fe³⁺ to Fe²⁺
- C Cl⁻ will reduce MnO4
- D HCl is a strong acid

(Total 1 mark)



Which one of the following can act as an oxidising agent but not as a reducing agent?

- A CH₃CHO
- **B** Fe²⁺
- **C** I⁻
- D MnO4



Aqueous $C_2 O_4^{2-}$ ions react with MnO4 ions in acidic solution according to the equation

$$5 C_2 O_4^{2-} + 2 Mn O_4^{-} + 16 H^+ \rightarrow 2 Mn^{2+} + 10 CO_2 + 8 H_2 O_2$$

Under the same conditions Fe^{2+} ions also react with MnO4 ions. How many moles of MnO4 ions are required to react exactly with one mole of $Fe(C_2O_4).2H_2O$?

- **A** 0.4
- **B** 0.6
- **C** 2.5
- **D** 7.5

(Total 1 mark)

Which one of the following statements about the reaction below is **false**?

 $[Cu(H_2O)_6]^{2+} + EDTA^{4-} \Longrightarrow [Cu(EDTA)]^{2-} + 6H_2O$

- **A** $[Cu(EDTA)]^{2-}$ is a more stable complex than $[Cu(H_2O)_6]^{2+}$
- **B** Both $[Cu(H_2O)_6]^{2+}$ and $[Cu(EDTA)]^{2-}$ are octahedral complexes.
- **C** There is an increase in entropy when the reaction occurs.
- **D** There is a redox reaction.

(Total 1 mark)

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Which one of the following would **not** react with aqueous silver nitrate to produce a precipitate that is soluble in concentrated aqueous ammonia?

- A CaBr₂
- **B** [COCl₄]²⁻
- C (CH₃)₄N⁺I⁻
- D CH₃COCI

(Total 1 mark)

4 Which one of the following would **not** reduce an acidified aqueous solution of potassium dichromate(VI)?

- A CH₃COOH
- **B** Zn
- C CH₃CHO
- D Fe²⁺(aq)

(Total 1 mark)

15 In the table below, which one of the following complex ions has a correct shape, co-ordination number and oxidation state?

	Complex	Shape	Co-ordination number	Oxidation state of central cation
А	[Ag(CN)₂] [−]	Linear	2	-1
В	[CuCl₄]²⁻	Tetrahedral	4	-2
С	[Cr(C₂O₄)₃] ³⁻	Octahedral	3	+3
D	[Cu(NH ₃)₄(H ₂ O) ₂] ²⁺	Octahedral	6	+2



Which one of the following could **not** act as a ligand?

- **A F**[−]
- B CH₃CH₃
- C NH₂NH₂
- D CH₃OCH₃

(Total 1 mark)

Use your knowledge of the chemistry of transition metals to predict which of the following will convert $[Mn(H_2O)_6]^{2*}$ into MnO4

- A an acid and a reducing agent
- **B** an acid and an oxidising agent
- C an alkali and a reducing agent
- D an alkali and an oxidising agent

(Total 1 mark)



The vanadium does not have an oxidation state of +3 in

- **A** $[V(H_2O)_6]^{3+}$
- **B** $[V(C_2O_4)_3]^{3-}$
- $C [V(OH)_3(H_2O)_3]$
- $D [VCI_4]^{3-}$





A 0.0720 g sample of reducing agent **R** was dissolved in water and acidified with an excess of dilute H_2SO_4 . The resulting solution was found to react with exactly 18.0 cm³ of a 0.0200 mol dm⁻³ solution of KMnO₄.

In this reaction, 5 mol of **R** react with 3 mol of KMnO₄. The M_r of **R** is

- **A** 120
- **B** 167
- **C** 240
- **D** 333

(Total 1 mark)



In which one of the following reactions is a heterogeneous catalyst not used?

- $\mathbf{A} \qquad \mathsf{N}_2 + \mathsf{3H}_2 \rightarrow \mathsf{2NH}_3$
- **B** $CO + NO \rightarrow CO_2 + \frac{1}{2}N_2$
- \mathbf{C} $CO_2 + C \rightarrow 2CO$
- $\mathbf{D} \quad \mathrm{SO}_2 + \frac{1}{2}\mathrm{O}_2 \rightarrow \mathrm{SO}_3$

(Total 1 mark)



Which one of the following reactions in aqueous solution has the most positive change in entropy?

- **A** $[Cu(H_2O)_6]^{2+} + 4NH_3 \rightarrow [Cu(NH_3)_4(H_2O)_2]^{2+} + 4H_2O$
- $\mathbf{B} \qquad [\mathsf{Cu}(\mathsf{H}_2\mathsf{O})_6]^{2*} + 4\mathsf{Cl}^{-} \rightarrow [\mathsf{Cu}\mathsf{Cl}_4]^{2-} + 6\mathsf{H}_2\mathsf{O}$
- $\mathbf{C} \qquad [\mathsf{Cu}(\mathsf{H}_2\mathsf{O})_6]^{2+} + \mathsf{E}\mathsf{DT}\mathsf{A}^{4-} \rightarrow [\mathsf{Cu}(\mathsf{E}\mathsf{DT}\mathsf{A})]^{2-} + 6\mathsf{H}_2\mathsf{O}$
- $\mathbf{D} \qquad [\mathsf{Cu}(\mathsf{H}_2\mathsf{O})_6]^{2*} + 2\mathsf{H}_2\mathsf{N}\mathsf{C}\mathsf{H}_2\mathsf{C}\mathsf{H}_2\mathsf{N}\mathsf{H}_2 \rightarrow [\mathsf{Cu}(\mathsf{H}_2\mathsf{N}\mathsf{C}\mathsf{H}_2\mathsf{C}\mathsf{H}_2\mathsf{N}\mathsf{H}_2)_2(\mathsf{H}_2\mathsf{O})_2]^{2*} + 4\mathsf{H}_2\mathsf{O}$