

GCE

Chemistry A

H032/01: Breadth in chemistry

AS Level

Mark Scheme for June 2026

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This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

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MARKING INSTRUCTIONS

PREPARATION FOR MARKING
RM ASSESSOR

1. Make sure that you have accessed and completed the relevant training packages for on-screen marking: *RM Assessor Assessor Online Training*; *OCR Essential Guide to Marking*.
2. Make sure that you have read and understood the mark scheme and the question paper for this unit. These are posted on the RM Cambridge Assessment Support Portal <http://www.rm.com/support/ca>
3. Log-in to RM Assessor and mark the **required number** of practice responses (“scripts”) and the **number of required** standardisation responses.

MARKING

1. Mark strictly to the mark scheme.
2. Marks awarded must relate directly to the marking criteria.
3. The schedule of dates is very important. It is essential that you meet the RM Assessor 50% and 100% (traditional 40% Batch 1 and 100% Batch 2) deadlines. If you experience problems, you must contact your Team Leader (Supervisor) without delay.
4. If you are in any doubt about applying the mark scheme, consult your Team Leader by telephone or the RM Assessor messaging system, or by email.
5. **Crossed Out Responses**
Where a candidate has crossed out a response and provided a clear alternative then the crossed out response is not marked. Where no alternative response has been provided, examiners may give candidates the benefit of the doubt and mark the crossed out response where legible.

Rubric Error Responses – Optional Questions

Where candidates have a choice of question across a whole paper or a whole section and have provided more answers than required, then all responses are marked and the highest mark allowable within the rubric is given. Enter a mark for each question answered into RM assessor, which will select the highest mark from those awarded. (*The underlying assumption is that the candidate has penalised themselves by attempting more questions than necessary in the time allowed.*)

Multiple Choice Question Responses

When a multiple choice question has only a single, correct response and a candidate provides two responses (even if one of these responses is correct), then no mark should be awarded (as it is not possible to determine which was the first response selected by the candidate).

When a question requires candidates to select more than one option/multiple options, then local marking arrangements need to ensure consistency of approach.

Contradictory Responses

When a candidate provides contradictory responses, then no mark should be awarded, even if one of the answers is correct.

Short Answer Questions (requiring only a list by way of a response, usually worth only **one mark per response**)

Where candidates are required to provide a set number of short answer responses then only the set number of responses should be marked. The response space should be marked from left to right on each line and then line by line until the required number of responses have been considered. The remaining responses should not then be marked. Examiners will have to apply judgement as to whether a 'second response' on a line is a development of the 'first response', rather than a separate, discrete response. *(The underlying assumption is that the candidate is attempting to hedge their bets and therefore getting undue benefit rather than engaging with the question and giving the most relevant/correct responses.)*

Short Answer Questions (requiring a more developed response, worth **two or more marks**)

If the candidates are required to provide a description of, say, three items or factors and four items or factors are provided, then mark on a similar basis – that is downwards (as it is unlikely in this situation that a candidate will provide more than one response in each section of the response space.)

Longer Answer Questions (requiring a developed response)

Where candidates have provided two (or more) responses to a medium or high tariff question which only required a single (developed) response and not crossed out the first response, then only the first response should be marked. Examiners will need to apply professional judgement as to whether the second (or a subsequent) response is a 'new start' or simply a poorly expressed continuation of the first response.

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Always check the pages (and additional objects if present) at the end of the response in case any answers have been continued there. If the candidate has continued an answer there, then add a tick to confirm that the work has been seen.

- 6.
7. There is a NR (No Response) option. Award NR (No Response)
 - if there is nothing written at all in the answer space
 - OR if there is a comment which does not in any way relate to the question (e.g. 'can't do', 'don't know')
 - OR if there is a mark (e.g. a dash, a question mark) which isn't an attempt at the question.

Note: Award 0 marks – for an attempt that earns no credit (including copying out the question).

Team Leaders must confirm the correct use of the NR button with their markers before live marking commences and should check this when reviewing scripts.

8. The RM Assessor **comments box** is used by your team leader to explain the marking of the practice responses. Please refer to these comments when checking your practice responses. **Do not use the comments box for any other reason.**

If you have any questions or comments for your team leader, use the phone, the RM Assessor messaging system, or e-mail.

9. Assistant Examiners will send a brief report on the performance of candidates to their Team Leader (Supervisor) via email by the end of the marking period. The report should contain notes on particular strengths displayed as well as common errors or weaknesses. Constructive criticism of the question paper/mark scheme is also appreciated.

On additional pages, annotate using 

10. Annotations

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3

Annotation	Meaning
	Benefit of doubt not given
	Noted but no credit given
	Ignore
	Blank page

11. Subject Specific Marking Instructions

Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

Annotation	Meaning
/; OR	alternative and acceptable answers for the same marking point
□	Separates marking points
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
—	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

INTRODUCTION

Your first task as an Examiner is to become thoroughly familiar with the material on which the examination depends. This material includes:

- the specification, especially the assessment objectives
- the question paper
- the mark scheme.

You should ensure that you have copies of these materials.

You should ensure also that you are familiar with the administrative procedures related to the marking process. These are set out in the OCR booklet **Instructions for Examiners**. If you are examining for the first time, please read carefully **Appendix 5 Introduction to Script Marking: Notes for New Examiners**.

Please ask for help or guidance whenever you need it. Your first point of contact is your Team Leader.

H032/01 Mark Scheme
SECTION A

June 2026

Question	Answer	Marks	Guidance
1	C	1	
2	B	1	
3	A	1	
4	B	1	
5	B	1	
6	C	1	ALLOW 7
7	C	1	
8	B	1	
9	D	1	
10	C	1	
11	A	1	
12	A	1	
13	A	1	
14	B	1	
15	C	1	
16	A	1	
17	A	1	
18	D	1	
19	D	1	
20	C	1	
	Total	20	

H032/01 Mark Scheme
SECTION B

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Question			Answer	Marks	Guidance										
21	(a)		$(1s^2)2s^22p^63s^23p^63d^84s^2$ ☐ $(1s^2)2s^22p^63s^23p^6$	2	ALLOW $4s^2$ before $3d^8$ i.e. $(1s^2)2s^22p^63s^23p^64s^2 3d^8$ ALLOW upper case D, etc and subscripts, e.g. $4S_23D_1$ ALLOW $4s^0$ IGNORE $[Ar]3d^8 4s^2$ IGNORE $[Ar]$										
21	(b)		<table><tr><td>Shell</td><td>1st</td><td>2nd</td><td>3rd</td><td>4th</td></tr><tr><td>Number of orbitals</td><td>1</td><td>4</td><td>9</td><td>16</td></tr></table> Requires all 4 numbers to be correct ☐	Shell	1st	2nd	3rd	4th	Number of orbitals	1	4	9	16	1	
Shell	1st	2nd	3rd	4th											
Number of orbitals	1	4	9	16											
21	(c)	(i)	Group 15 AND p-block	1	ALLOW Group 5 IGNORE superscripts/numbers given on block (e.g. p^5) if the letter is clear										
21	(c)	(ii)	$Na_3As + 3HCl \rightarrow AsH_3 + 3NaCl$ AsH_3 Correct balanced equation	2	ALLOW multiples ALLOW H_3As IGNORE state symbols No ECF from any incorrect formula										

Question		Answer	Marks	Guidance
21	(d)	HI / I has more electrons (than HBr / Br) □ In HI, stronger/more induced dipole–dipole interactions OR stronger/more London forces (than HBr) OR more energy required to overcome induced dipole–dipole interactions / London forces	2	Both answers need to be a comparison ORA throughout ALLOW names if no formulae given ALLOW larger electron cloud ALLOW more electron shells IGNORE just ‘more shells’ OR more (electron) shielding IGNORE HI molecules are bigger or have greater Mr. IGNORE comments about nuclear attraction, ionisation energy, etc. IGNORE comments about electronegativity ALLOW instantaneous/temporary–induced dipole interactions ALLOW London dispersion forces IGNORE van der Waals’ / vdW / IDD / IDID IGNORE just ‘bonds’ DO NOT ALLOW hydrogen bonds DO NOT ALLOW response that implies covalent or ionic bonds or permanent dipole-dipole interactions breaking ALLOW first marking point ONLY if answer states ‘down Group 7 there are more electrons and more London forces’ (unless directly linked to H-Br/HI for second marking point)

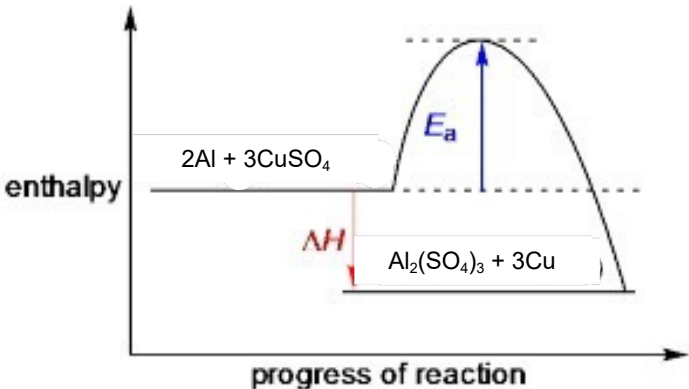
Question			Answer	Mark	Guidance
22	(a)	(i)	$3\text{Mg(s)} + 2\text{H}_3\text{PO}_4\text{(aq)} \rightarrow \text{Mg}_3(\text{PO}_4)_2\text{(aq)} + 3\text{H}_2\text{(g)}$ ☐	1	ALLOW multiples ALLOW 1 for $\text{Mg}_3(\text{PO}_4)_2$
22	(a)	(ii)	Element Oxidised: Magnesium/Mg $0 \rightarrow +2$ ☐ Element Reduced: Hydrogen/H $+1 \rightarrow 0$ IGNORE oxidation numbers shown below/above equation (treat as rough working) BUT If no oxidation numbers in explanation, look at equation for oxidation numbers	2	ALLOW 1 out of 2 redox marks if oxidation number changes are BOTH correct BUT reduction/oxidation is incorrectly assigned, i.e written on the wrong line. General: ALLOW number before sign in ox no, e.g. 1– for –1 IGNORE '1', '2' (signs required) ALLOW H2 for hydrogen IGNORE references to electron loss/gain (even if wrong)

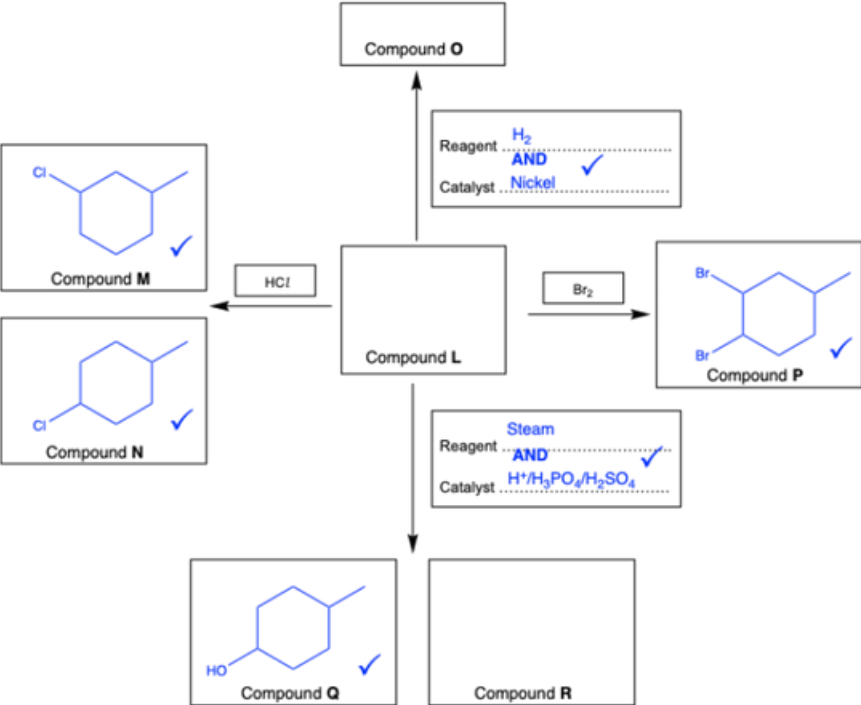
Question		Answer	Mark	Guidance
22	(b)	$\text{Mg}^+(\text{g}) \rightarrow \text{Mg}^{2+}(\text{g}) + \text{e}^-$ <i>Equation correct</i> <i>Correct state symbols</i> ☐ ----- Ionic radius $\text{Mg}^{(+)}$ has smaller (ionic) radius OR Magnesium (ion) has fewer shells OR $\text{Mg}^{(+)}$ has less/decreased shielding Attraction Nuclear attraction is greater OR (outer) electron(s) experience greater nuclear attraction	4	ALLOW $\text{Mg}^+(\text{g}) - \text{e}^- \rightarrow \text{Mg}^{2+}(\text{g})$ for 2 marks The second mark is dependent upon the first mark except for the following close attempts: ALLOW one mark for the following for state symbols $\text{Mg}(\text{g}) \rightarrow \text{Mg}^{2+}(\text{g}) + 2\text{e}^-$ $\text{Mg}^+(\text{g}) + \text{e}^- \rightarrow \text{Mg}^{2+}(\text{g}) + 2\text{e}^-$ $\text{Mg}(\text{g}) \rightarrow \text{Mg}^+(\text{g}) + \text{e}^-$ ALLOW e for electron (i.e. charge omitted) IGNORE states on the electron ALLOW ORA throughout (but assume 'it' refers to Mg^+) IGNORE atomic IGNORE more orbitals OR more sub-shells <i>Alternative must refer to shells</i> ALLOW Energy levels for shells ALLOW both electron configurations for 'fewer shells' ALLOW removes electron from shell closer to the nucleus ALLOW more electron repulsion between shells IGNORE just 'shielding' (<i>more/greater needed</i>) ALLOW 'pull' for attraction IGNORE 'electrons less tightly held' IGNORE 'nuclear charge' for 'nuclear attraction' IGNORE comments about reactivity and energy

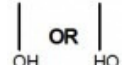
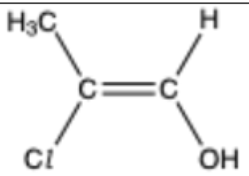
Question			Answer	Mark	Guidance																
23	(a)	(i)	<table><tr><td>Trial</td><td>1</td><td>2</td><td>3</td></tr><tr><td>0.00</td><td>10.00</td><td>10.00</td><td>5.00</td></tr><tr><td>32.50</td><td>41.60</td><td>42.10</td><td>36.50</td></tr><tr><td>32.50</td><td>31.60</td><td>32.10</td><td>31.50</td></tr></table> All 4 titres correct to 2 DP	Trial	1	2	3	0.00	10.00	10.00	5.00	32.50	41.60	42.10	36.50	32.50	31.60	32.10	31.50	1	DO NOT ALLOW 1DP, e.g. 32.1 instead of 32.10
Trial	1	2	3																		
0.00	10.00	10.00	5.00																		
32.50	41.60	42.10	36.50																		
32.50	31.60	32.10	31.50																		
23	(a)	(ii)	Meantitre = $\frac{31.60 + 31.50}{2}$ = 31.55 (cm³) <i>i.e. using concordant (consistent) titre</i>	1	DO NOT ALLOW mean of all three titres, i.e. = $\frac{31.60 + 32.10 + 31.50}{3}$ = 31.7(3) ALLOW ECF from incorrect concordant titres from 22a(i); ALLOW ECF mean titres to either 2 or 3 dp DONOTALLOW use of trial titre.																
23	(a)	(iii)	FIRST CHECK THE ANSWER ON ANSWER LINE IF answer = 1.14×10^{-1} (mol dm⁻³) award 3 marks ----- $n(\text{Ba}(\text{OH})_2) = 0.12 \times 0.015$ = 1.80 10^{-3} (mol) $n(\text{HNO}_3) = 2 \times n(\text{Ba}(\text{OH})_2) = \textbf{3.60}$ 10^{-3} (mol) concentration of $\text{HNO}_3(\text{aq})$ = $\frac{3.60 \times 10^{-3}}{0.03155}$ = 1.14 10^{-1} (mol dm⁻³) 3SF AND Standard form required	3	Working to at least 3SF but IGNORE 'trailing zeroes' e.g. ALLOW 1.8 for 1.80 ----- ALLOW ECF from incorrect mean titre in a(ii) ALLOW ECF from 2 incorrect $n(\text{Ba}(\text{OH})_2)$ ALLOW ECF from incorrect $n(\text{HNO}_3)$ Common errors 5.71×10^{-2} (no x2) → 2 marks 2.85×10^{-2} (÷2 instead of x 2) → 2 marks 0.114 (no standard form) → 2 marks																

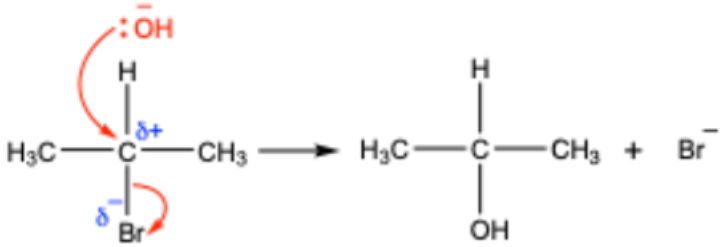
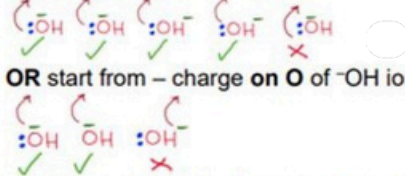
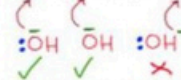
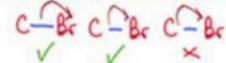
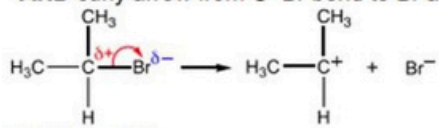
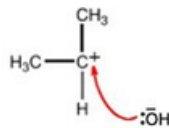
Question		Answer	Mark	Guidance
23	(b)	Equal/same number of moles of MgCO_3 used (forming the same volume of gas) HNO_3 is a strong acid/completely dissociates/completely ionises AND CH_3COOH is a weak acid/partially dissociates/partially ionises Greater H^+ concentration in HNO_3 AND More frequent collisions / faster (rate of) reaction	3	ALLOW names if no formulae given ALLOW same amount of MgCO_3 ALLOW same number of carbonate/CO_3^{2-} ions IGNORE same mass of MgCO (<i>in question</i>) IGNORE same concentration / moles of acids IGNORE limiting reactant ALLOW HNO_3 is a stronger acid/dissociates more than ethanoic acid. ALLOW <i>idea of</i> more H^+ e.g. 'releases all H^+ ions' / 'more H^+ ions dissociate from acid' ALLOW ORA

Question			Answer	Mark	Guidance
24	(a)	(i)	FIRST, CHECK THE ANSWER ON ANSWER LINE If = -508 (kJ mol⁻¹) award 4 marks ----- Energy released in J OR kJ =50 4.18 40.5 = 8464.5 (J) OR 8.4645 (kJ) Correctly calculates $n(\text{CuSO}_4)$ = $\frac{50 \times 1}{1000}$ = 0.0500 (mol) Energy change per mole CuSO₄ = $\frac{8464.5}{0.0500}$ OR $\frac{8.4645}{0.0500}$ = ±169290 (J mol⁻¹) OR ±169.29 (kJ mol⁻¹) Value of ΔH in kJ mol⁻¹ to 3 SF = -508 (kJ mol⁻¹) AND 3SF AND -sign	4	FULL ANNOTATIONS MUST BE USED ----- ALLOW ECF throughout ALLOW -507 (kJ mol⁻¹) from $n(\text{Al}_2(\text{SO}_4)_3) = 0.0167$ (3 SF) Common Errors for 3 marks ±169 (no scaling) +508 (wrong sign) ±507.87 (not 3SF) ±56.4 (divide by 3, not multiply) -510 (use of 8.5 kJ, use of 2SF) -3930 (use of 313.5K) -81.1 (use of 7.98g [0.05 x Mr] for mass) – 627 (use of 7.98g and 313.5 K) Common Errors for 2 marks ±27.0 (use of 7.98g [0.05 x Mr] for mass; no scaling) +510 (use of 8.5 kJ, use of 2SF and sign)
24	(a)	(ii)	Any ONE from: 1 MARK ONLY ☐ • Insulate beaker • Change beaker for polystyrene cup • Add lid to beaker	1	ALLOW any answer which counteracts heat loss e.g. thermos flask ALLOW use of standard conditions ALLOW use a cooling curve ALLOW measure mass of solution ALLOW use burette to measure volume ALLOW use of more accurate/precise/digital thermometer IGNORE use of copper beaker

Question	Answer	Mark	Guidance
24 (b)	 2Al + 3CuSO₄ on LHS AND Al₂(SO₄)₃ + 3Cu on RHS AND ΔH labelled with product line below reactant line AND Arrow downwards Ea (independent of ΔH) curve with arrow from reactants to top of curve AND Ea labelled ☐	2	Mark ALL exothermic profiles without reference to 24(a)(i) If endothermic profile seen, check answer line in 24(a)(i) and ALLOW endothermic profile as ECF from a positive energy calculation/+ΔH value for two marks State symbols NOT required ΔH DONOT ALLOW – ΔH DONOT ALLOW doubleheaded arrow on ΔH ALLOW ΔH arrow even with small gap at the top and bottom, i.e. line does not quite reach reactant or product line. ECF on endothermic profile: ΔH labelled with product line above reactant line AND arrow upwards Ea ALLOW no arrowhead or arrowheads at both ends of line Ea BUT DO NOT ALLOW arrowhead down Ea line must reach maximum (or near to maximum) on curve For Ea, ALLOW AE OR A_E OR Eact OR suitable alternatives ALLOW ECF marks for Ea for correctly labelled endothermic diagram from a –ΔH value (from 24(a)(i))

Question	Answer	Mark	Guidance
25 (a) (i)	4-methylcyclohex(-1-)ene ☐	1	ALLOW two words: 4-methylcyclohexene OR three words: 4-methyl cyclo hexene DO NOT ALLOW meth OR methy IGNORE lack of hyphens, extra hyphens, full stops instead of commas, extra spaces
25 (a) (ii)		6	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous DO NOT ALLOW structure if H(s) are missing from ONE structural/displayed formula... BUT ALLOW any further omissions as ECF For M and N, DO NOT ALLOW structure if H is added to a carbon in the skeletal formula BUT ALLOW ECF for the other structure ALLOW names of reagents and catalysts ALLOW reagent and catalyst on either answer line DO NOT ALLOW other additional reagents ALLOW Pt OR Pd for Ni ALLOW in either order (M and N) ALLOW H₂O(g) for steam (g) OR temperature >100°C required ALLOW 'acid' for H⁺ BUT DO NOT ALLOW HCl ALLOW small slip in acid formula e.g. phosphoric acid as H₂PO₃, etc ALLOW any vertical bond to the OH

Question			Answer	Mark	Guidance
					 DO NOT ALLOW OH– e.g. where OH– drawn as diagonal in Compound Q Modified braille paper: V = Q Vi = M Vii = N Viii = P
25	(a)	(iii)	Mechanism: Electrophilic addition Bond fission: Heterolytic	2	IGNORE just ‘addition’ IGNORE just ‘electrophilic’
25	(a)	(iv)	Observation: turns (orange to) colourless OR decolorises AND Functional group: alkene/ C=C <i>Take care – only one mark for both points</i>	1	IGNORE colour change IGNORE clear
25	(b)			1	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous IGNORE connectivity ALLOW <i>E</i> isomer Modified braille paper: Correct answer is ‘A’.

Question	Answer	Mark	Guidance
26 (a) (i)	An electron pair donor ☐	1	ALLOW lone pair donor
(a) (ii)	FULL ANNOTATIONS MUST BE USED Mechanism Curly arrow from OH⁻ to C atom of C-Br bond in 2-bromopropane Dipole shown on C-Br bond, C⁺ and Br⁻, AND curly arrow from C-Br bond to Br atom  IGNORE presence of Na⁺ but OH⁻ needed i.e. Na⁺OH⁻ can be allowed if criteria met NOTE curly arrows can be straight or snake-like but double-headed or half-headed arrows NOT	3	1st curly arrow must - go to the C of C-Br AND - start from, OR be traced back to any point across width of lone pair on O of OH⁻  OR start from - charge on O of OH⁻ ion  (Lone pair NOT needed if curly arrow shown from O-) 2nd curly arrow must start from, OR be traced back to, any part of C-Br bond and go to Br  ALLOW S_N1 mechanism for 2 curly arrow marks First mark Dipole shown on C-Br bond, C⁺ and Br⁻, AND curly arrow from C-Br bond to Br atom ✓  Second mark Curly arrow from OH⁻ AND to correct carbocation ✓ Curly arrow must come from lone pair on O of HO⁻ OR OH⁻ OR from minus on O of HO⁻ ion (no need to show lone pair if curly arrow came from negative charge) ✓ 

Question	Answer	Mark	Guidance
	Products 1 mark correct organic product AND Br⁻ ☐ $\begin{array}{c} \text{H} \\ \\ \text{H}_3\text{C}-\text{C}-\text{CH}_3 \\ \\ \text{OH} \end{array} + \text{Br}^-$ IGNORE presence of Na⁺ but Br⁻ needed i.e. Na⁺Br⁻ can be allowed BUT NaBr does NOT show Br⁻		FOR ORGANIC PRODUCT ALLOW any combination of skeletal formula as long as unambiguous IGNORE connectivity throughout OR structural OR displayed Modified braille paper: Students will respond descriptively. No expectation that they 'draw' the mechanism. Use answer and guidance wording for marking.
26 (b)	Reagents: Water OR aqueous/(aq) seen anywhere ☐ AgNO₃/silver nitrate AND ethanol ☐	5	FULL ANNOTATIONS MUST BE USED ALLOW formulae if no name given ALLOW name if no formulae given IGNORE any reference to hydroxides IGNORE any reference to NH₃/ammonia DO NOT ALLOW H₂O(g)/steam DO NOT ALLOW 'acidified' or mention of acids

Question	Answer	Mark	Guidance
	Observations Any two from: 1 MARK ONLY • white (precipitate) • cream (precipitate) • yellow (precipitate) Rates of hydrolysis (fastest to slowest) (2-)iodopropane/C-I, (2-)bromopropane/C-Br, (2-)chloropropane/C-Cl Explanation C-I bond is weaker than C-Br bond AND C-Br bond is weaker than C-Cl bond OR C-I bond has a lower bond enthalpy than C-Br bond AND C-Br bond has a lower bond enthalpy than C-Cl bond □		DO NOT ALLOW any other colours IGNORE reactivity of halogens ALLOW unambiguous comparison of two haloalkanes e.g. (2-)iodopropane is fastest AND (2-)chloropropane is slowest ALLOW rate of formation of AgX e.g. AgI precipitate forms fastest, AgCl precipitate forms slowest ALLOW ORA There must be a comparison between C-Cl, C-Br and C-I bonds but ALLOW unambiguous comparison of two bonds e.g. C-I bond is the weakest AND C-Cl is the strongest bond IGNORE references to bond length, polarity and electronegativity

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