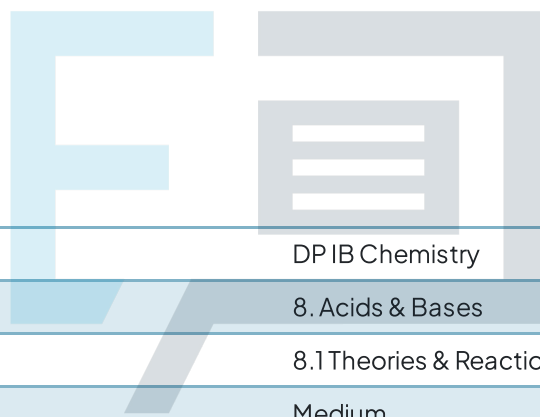




# 8.1 Theories & Reactions of Acids & Bases

## Mark Schemes



Course	DP IB Chemistry
Section	8. Acids & Bases
Topic	8.1 Theories & Reactions of Acids & Bases
Difficulty	Medium

# Exam Papers Practice

To be used by all students preparing for DP IB Chemistry SL  
Students of other boards may also find this useful



1

The correct answer is **A** because:

- Acids react with hydrogen carbonates to form a salt, water and carbon dioxide
- Acids react with magnesium to form a salt and hydrogen
- Copper is too low in the reactivity series to react with dilute acids

#### Extra info:

Although metals generally react with most acids, copper is an exception. Other familiar metals that would also **not** react with acids are silver, mercury and gold

2

The correct answer is **D** because:

- A Brønsted–Lowry base is defined as a proton ( $\text{H}^+$ ) acceptor
- The  $\text{HNO}_3$  gains a proton to become  $\text{H}_2\text{NO}_3^+$  so it must be acting as a base
- In the reverse reaction, the  $\text{HSO}_4^-$  gains a proton to become  $\text{H}_2\text{SO}_4$  so it must also be acting as a base

**A** is incorrect as  $\text{H}_2\text{NO}_3^+$  acts as Brønsted–Lowry acid because it is donating a proton to  $\text{HSO}_4^-$

**B** is incorrect as  $\text{H}_2\text{NO}_3^+$  acts as Brønsted–Lowry acid because it is donating a proton to  $\text{HSO}_4^-$

**C** is incorrect as  $\text{H}_2\text{SO}_4$  acts as Brønsted–Lowry acid because it is donating a proton to  $\text{HNO}_3$



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The correct answer is **D** because:

- A proton is transferred from the HCl to  $\text{CH}_3\text{COOH}$  making the HCl an acid and  $\text{CH}_3\text{COOH}$  a base
- In the back reaction, a proton is transferred from  $\text{CH}_3\text{COOH}_2^+$  to  $\text{Cl}^-$  making the  $\text{CH}_3\text{COOH}_2^+$  an acid and  $\text{Cl}^-$  a base

**A** is incorrect as both the acid and base in the reactants and products are the wrong way around

**B** is incorrect as the acid and base in the reactants are the wrong way around

**C** is incorrect as the acid and base in the products are the wrong way around

4

The correct answer is **C** because:

- You are told that perbromic acid is a strong acid, so by definition it must be fully dissociated in solution

**A** is incorrect as this would be a feature of weak acids, not strong ones

**B** is incorrect as a strong acid would react with weak or strong bases

**D** is incorrect as a strong or weak acid would have a pH below 7



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The correct answer is **C** because:

- To form a conjugate base, a species must lose a proton
- When  $\text{K}_2\text{HPO}_4$  loses a proton it forms  $\text{K}_2\text{PO}_4^-$
- However, this does not seem to be one of the answers
  - **Remember:** The potassium ion is a spectator ion and positively charged so the anion corresponds to  $\text{PO}_4^{3-}$  ( $\text{K}_2\text{PO}_4^- = 2\text{K}^+ + \text{PO}_4^{3-}$ )

**A** is incorrect as this is the conjugate acid of the hydrogen phosphate ion  $\text{K}_2\text{HPO}_4 + \text{H}^+ \rightarrow \text{KH}_2\text{PO}_4$  ( $\text{H}_2\text{PO}_4^-$  is the anion)

**B** is incorrect as this ion does not make sense

**D** is incorrect as this is the full formula of the conjugate acid of the hydrogen phosphate ion

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The correct answer is **D** because:

- To be a Brønsted-Lowry acid, a species must lose a proton
- In the forward reaction,  $\text{H}_3\text{NSO}_3$ , loses protons and in the backward reaction each  $\text{NH}_4^+$  loses a proton

**A** is incorrect as this is a Brønsted-Lowry acid-base pair

**B** is incorrect as this is a Brønsted-Lowry acid and a Brønsted-Lowry base, respectively

**C** is incorrect as this is a Brønsted-Lowry acid-base pair



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The correct answer is **D** because:

- Calcium oxide is classified as a base
- When an acid reacts with a base a salt and water are formed
- Salts of hydrochloric acid are chlorides

**A** is incorrect as there is no gas given off in the reaction and water is a missing product

**B** is incorrect as there is no gas given off in the reaction

**C** is incorrect as calcium and hydrogen are incorrect products

8

The correct answer is **D** because:

- The correctly balanced equation is:  
$$2\text{HNO}_3(\text{aq}) + \text{Mg}(\text{HCO}_3)_2(\text{s}) \rightarrow \text{Mg}(\text{NO}_3)_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l}) + 2\text{CO}_2(\text{g})$$
- The sum of the coefficients =  $2 + 1 + 1 + 2 + 2 = 8$

**A, B & C** are incorrect as they do not have the right sum of coefficients

It is a favourite examiner's trick to ask for the sum of the coefficients knowing that some candidates will forget that the absence of a number in front of a formula means the coefficient is 1. Don't get caught out!

9

The correct answer is **A** because:

- Calcium oxide is classified as a base
- Only oxides that are acidic can react with calcium oxide
- Sulfur dioxide and nitrogen dioxide are acidic oxides

**B** is incorrect as nitrogen dioxide reacts with calcium oxide, but potassium oxide does not

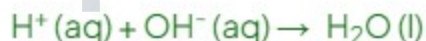
**C** is incorrect as sulfur dioxide and nitrogen dioxide react with calcium oxide, but potassium oxide does not

**D** is incorrect as potassium oxide does not react with calcium oxide

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The correct answer is **B** because:

- Neutralisation reactions are always exothermic as there is net bond formation when hydrogen ions react with hydroxide ions to form water:



- Bond formation is always an exothermic process

**A** is incorrect as metal displacement reactions are generally exothermic, but they can be either

**C** is incorrect as combustion is always an exothermic reaction

**D** is incorrect as melting is an endothermic process as it is a bond breaking process