

Friday 10 June 2016 – Morning

**GCSE GATEWAY SCIENCE
CHEMISTRY B**

B741/01 Chemistry modules C1, C2, C3 (Foundation Tier)

Candidates answer on the Question Paper.
A calculator may be used for this paper.

OCR supplied materials:
None

Other materials required:

- Pencil
- Ruler (cm/mm)

Duration: 1 hour 15 minutes



Candidate forename		Candidate surname	
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Centre number						Candidate number				
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INSTRUCTIONS TO CANDIDATES

- Write your name, centre number and candidate number in the boxes above. Please write clearly and in capital letters.
- Use black ink. HB pencil may be used for graphs and diagrams only.
- Answer **all** the questions.
- Read each question carefully. Make sure you know what you have to do before starting your answer.
- Write your answer to each question in the space provided. If additional space is required, you should use the lined page(s) at the end of this booklet. The question number(s) must be clearly shown.
- Do **not** write in the bar codes.

INFORMATION FOR CANDIDATES

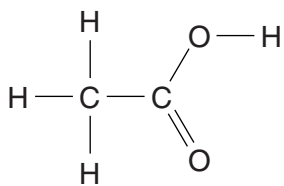
- The quality of written communication is assessed in questions marked with a pencil (✎).
- The Periodic Table can be found on the back page.
- The number of marks is given in brackets [] at the end of each question or part question.
- The total number of marks for this paper is **75**.
- This document consists of **28** pages. Any blank pages are indicated.

Answer **all** the questions.

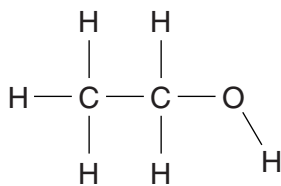
SECTION A – Module C1

1 This question is about carbon compounds.

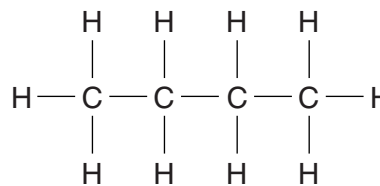
Look at the displayed formulas of some compounds.



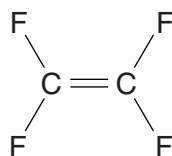
A



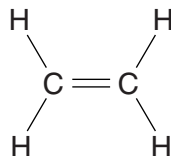
B



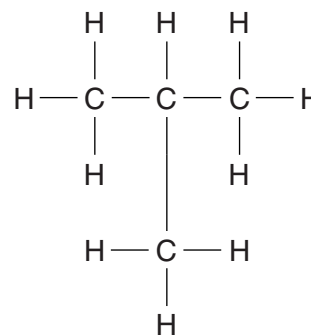
C



D



E



F

(a) Which compound has a molecule with **8 atoms**?

Choose from **A, B, C, D, E** or **F**.

.....

[1]

(b) Which compound is an **alkene**?

Choose from **A, B, C, D, E** or **F**.

.....

[1]

(c) Which **two** compounds have the same **molecular** formula?

Choose from **A, B, C, D, E** and **F**.

..... and

[1]

(d) Compound **D** is called tetrafluoroethene.

It is a monomer and can be made into a polymer.

What is the name of this polymer?

..... [1]

2 Esters can be used in nail varnish remover.

Look at the table of information about some esters.

Ester	Molecular formula of ester	Melting point in °C	Boiling point in °C	How well it dissolves in water (1 = poor 10 = good)
methyl ethanoate	$C_3H_6O_2$	-98	57	1
ethyl ethanoate	$C_4H_8O_2$	-84	77	8
propyl ethanoate	$C_5H_{10}O_2$	-95	102	2
butyl ethanoate	$C_6H_{12}O_2$	-77	127	1
pentyl ethanoate		-71	149	1

(a) Esters are **not** hydrocarbons.

Explain why using information from the molecular formulas.

.....
 [1]

(b) Pentyl ethanoate has **seven** carbon atoms in its molecule.

Deduce the **molecular formula** for pentyl ethanoate.

..... [1]

(c) The solvent in a nail varnish remover needs to have these properties

- liquid at room temperature, 25 °C
- evaporates easily
- fairly soluble in water.

Which ester would be **most** suitable for use as a nail varnish remover?

Explain your answer.

.....

 [3]

[Total: 5]

Turn over

4 The Olympic flame for the London Olympics burned a fuel. This fuel was a gas.



(a) The designer of the Olympic flame had to decide which fuel to use.

He could not decide whether to use bio gas or natural gas.

One factor he considered was the energy value of the two fuels.

Write about **three** other factors he had to consider.

.....
.....
.....
.....
.....
..... [3]

(b) The designer decided to use natural gas.

The Olympic flame was yellow in colour and could be seen easily.

Natural gas normally burns with a blue flame.

Suggest **two** reasons why the Olympic flame was yellow and not blue.

.....
.....
..... [2]

[Total: 5]

Turn over

5 This question is about different paints.

Look at the table. It shows the percentage by mass of each ingredient in four paints.

Ingredient	Percentage by mass in each paint			
	Paint A	Paint B	Paint C	Paint D
solvent	32	25	55	20
oil	0	25	0	25
pigment	24	10	5	30
bonding medium	30	30	28	23
other additives	14	10	12	2

(a) Look at the column for paint D.

Parminder wants to present the data about the ingredients in paint D.

Which will be the **best** way for her to present this data?

Choose from

bar chart

line graph

pie chart

scatter graph

answer [1]

(b) Parminder wants to show the percentage of solvent in each of the four paints.

Which will be the **best** way for her to present this data?

Choose from

bar chart

line graph

pie chart

scatter graph

answer [1]

(c) Which paint is likely to be the easiest to spread?

.....

Explain your answer.

.....

..... [1]

(d) Paint **B** contains a **thermochromic** pigment.

Paint **D** contains a **phosphorescent** pigment.

Write about the differences between a thermochromic pigment and a phosphorescent pigment.

.....

.....

.....

..... [2]

[Total: 5]

Question 6 begins on page 8

SECTION B – Module C2

6 This question is about metals and alloys.

(a) Look at the list of materials.

amalgam

bronze

copper

mercury

solder

Some of the materials are **alloys**.

Some of the materials are **metallic elements**.

Finish the table. Put each material in the correct column.

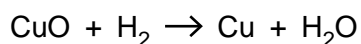
One material has been done for you.

Alloy	Metallic element
	copper

[2]

(b) Copper can be made from copper oxide, CuO, using hydrogen gas.

Look at the equation.



This reaction is an example of **reduction**.

Copper oxide is reduced.

Explain how you can tell from the equation.

.....

.....

..... [1]

(c) Look at the table. It shows the properties of some alloys.

Alloy	Density in g/cm ³	Relative strength	Relative cost
duralumin	2.8	high	high
brass	8.4	high	medium
steel	7.8	very high	low



Which alloy is most suitable for making aeroplane wings?

Use information from the table to give **two** reasons for your choice.

.....

.....

.....

.....

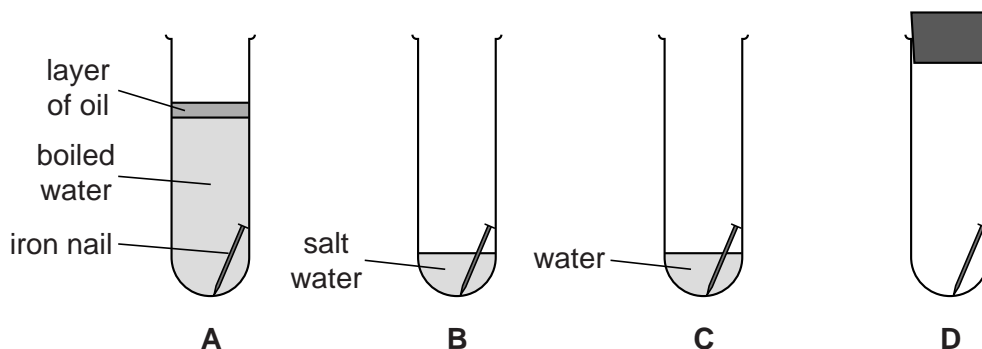
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..... [2]

[Total: 5]

7 Robert investigates the rusting of iron.

Look at the diagram. It shows the apparatus he uses.



Substances in contact with iron	boiled water with no dissolved air	air and salt water	air and water	air with no water
Results	not rusty after 10 days	rusty after 24 hours	rusty after 5 days	not rusty after 10 days

(a) (i) Robert thinks that **both** air and water are needed for iron to rust.

How do you know that he is right?

Use the diagram to help you.

.....

 [2]

(ii) Which substance in Robert's investigation **speeds up** rusting?

..... [1]

(b) Aluminium, Al , does not corrode in air.

This is because aluminium reacts with oxygen, O_2 .

A protective layer of aluminium oxide, Al_2O_3 , is made.

Write a **balanced symbol** equation for this reaction.

..... [2]

[Total: 5]

11
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Question 8 begins on page 12
PLEASE DO NOT WRITE ON THIS PAGE

8 This question is about making ammonia by the Haber process.

(a) Nitrogen and hydrogen react to make ammonia.

(i) The hydrogen gas used to make ammonia comes from natural gas.

Where does the **nitrogen gas** come from?

Choose from the list.

air

limestone

salt

water

..... [1]

(ii) Nitrogen and hydrogen react to make ammonia.

Not all of the nitrogen and hydrogen reacts.

What happens to the unreacted nitrogen and hydrogen?

..... [1]

- 9 In 2012 bad weather destroyed farmers' crops in Burma.



The charity called Oxfam helped farmers after the disaster.

Oxfam gave the farmers

- seeds to plant new crops
- fertilisers.

(a) Why do farmers use fertilisers?

..... [1]

(b) Potassium sulfate, K_2SO_4 , is a fertiliser.

Potassium sulfate contains the **essential** element potassium, K.

Ammonium phosphate, $(NH_4)_3PO_4$, is another fertiliser.

Write down the names of the **two essential** elements in ammonium phosphate.

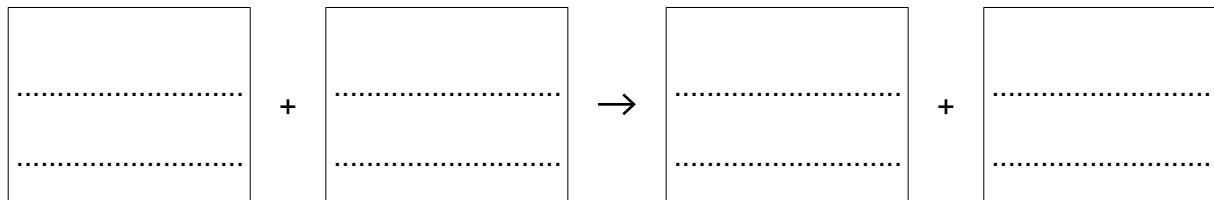
.....
 [2]

(c) Fertilisers are made by reacting an acid with an alkali.

This is a neutralisation reaction.

(i) Potassium hydroxide reacts with nitric acid.

Write a **word equation** for this reaction.



[2]

(ii) One way to find out the pH of a solution of fertiliser is by using a pH meter.

Write about one **other** way to find the pH.

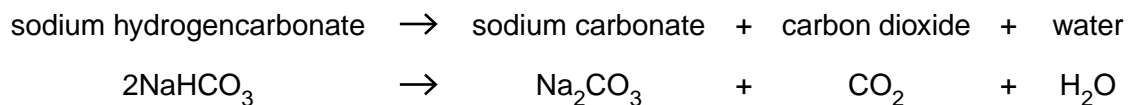
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..... [2]

[Total: 7]

Question 10 begins on page 16

SECTION C – Module C3

10 Sodium hydrogencarbonate decomposes when it is heated.



The table shows the relative formula masses, M_r , of the substances in the equation.

Substance	Relative formula mass
NaHCO ₃	84
Na ₂ CO ₃	106
CO ₂	44
H ₂ O	18

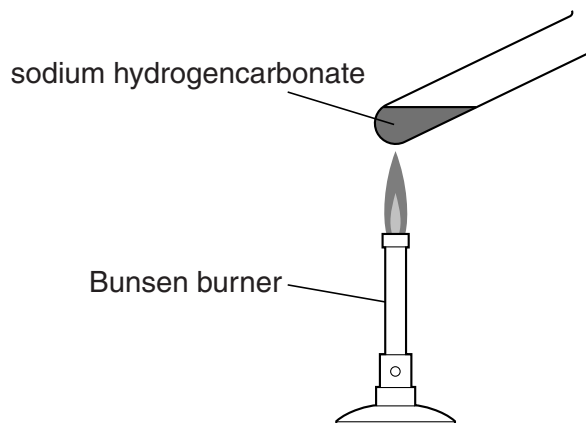
(a) Show that the relative formula mass of Na₂CO₃ is 106.

The relative atomic mass, A_r , of C = 12, O = 16 and of Na = 23.

..... [1]

(b) Zakia heats some sodium hydrogencarbonate.

Look at the apparatus she uses.



Zakia heats 0.84 g of solid sodium hydrogencarbonate.

When the reaction is complete the test tube contains 0.53 g of solid sodium carbonate.

(i) The mass of the solid in the test tube **decreases** when it is heated.

Explain why.

.....
 [1]

(ii) Zakia does another experiment.

This time she heats 8.4 g of solid sodium hydrogencarbonate.

Predict the mass of solid sodium carbonate made when the reaction is complete.

.....
 mass = g [1]

(iii) Zakia makes **less** solid sodium carbonate than she predicts.

Suggest **two** reasons why she makes less solid.

.....

 [2]

[Total: 5]

11 Phil is a research chemist.

He investigates a new pharmaceutical drug.

Phil extracts the drug from the leaves of a plant.

He purifies the drug and then checks to see if he has made a pure sample.

Phil uses two tests to check the purity of the drug

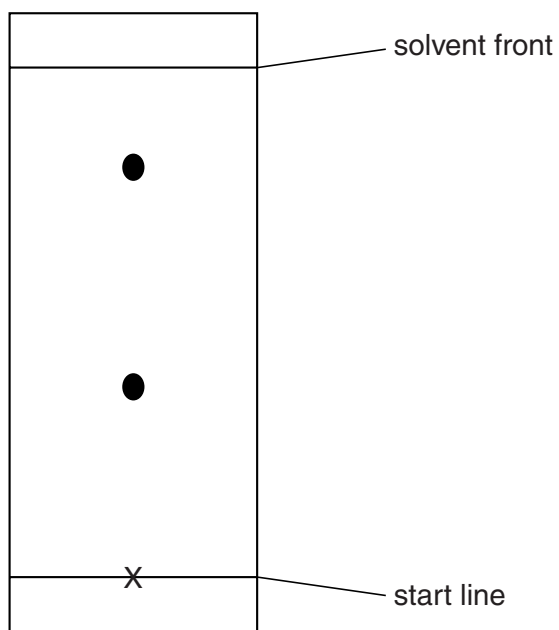
- melting point
- thin layer chromatography.

Look at the results of his tests.

Melting point

Substance	Melting point in °C
pure drug	175
sample of the drug obtained from plant	171 – 173

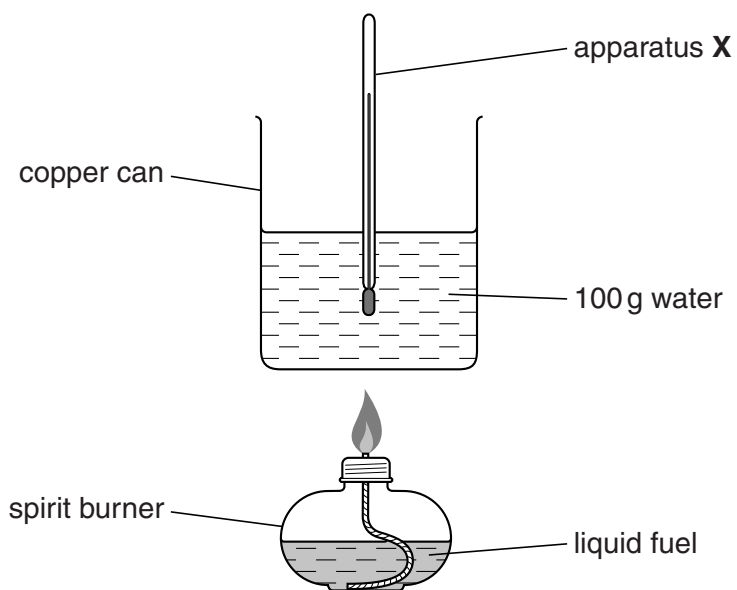
Thin layer chromatogram of sample of the drug obtained from the plant.



12 Zak compares different fuels.

He heats 100 g of water each time.

Look at the apparatus he uses.



(a) Look at the diagram.

What is the name of apparatus X?

..... [1]

(b) Zak uses four fuels.

Zak always burns the same mass of fuel.

Explain why.

.....
 [1]

(c) Look at Zak's table of results.

Fuel	Temperature of water at start in °C	Temperature of water after heating in °C	Temperature change of water in °C
A	22	45	23
B	22	48
C	21	48
D	17	47

Calculate the temperature change for each fuel. One has been done for you.

Use these results to decide which fuel releases the most heat energy.

.....
 [2]

(d) What is the name of the **type** of chemical reaction that gives out energy into the surroundings?

..... [1]

[Total: 5]

Question 13 begins on page 22

13 Zinc reacts with hydrochloric acid.

Hydrogen gas and zinc chloride are made.

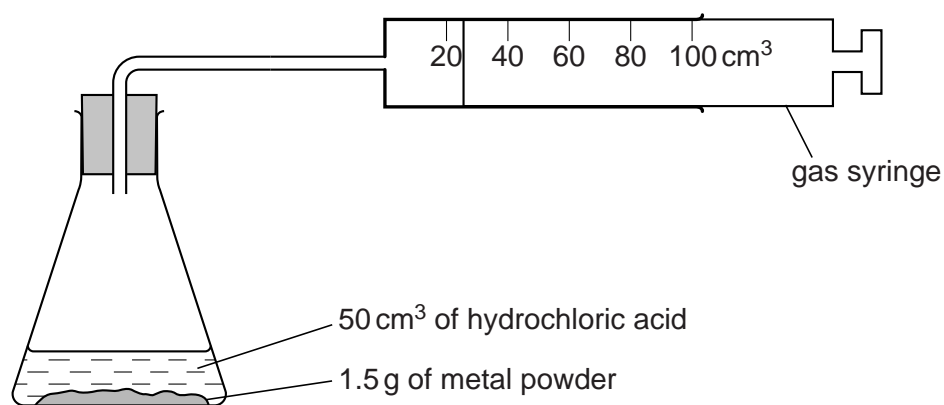
(a) Write the **word** equation for this reaction.

..... [1]

(b) Fatimah and Sam investigate the reaction between acid and metals.

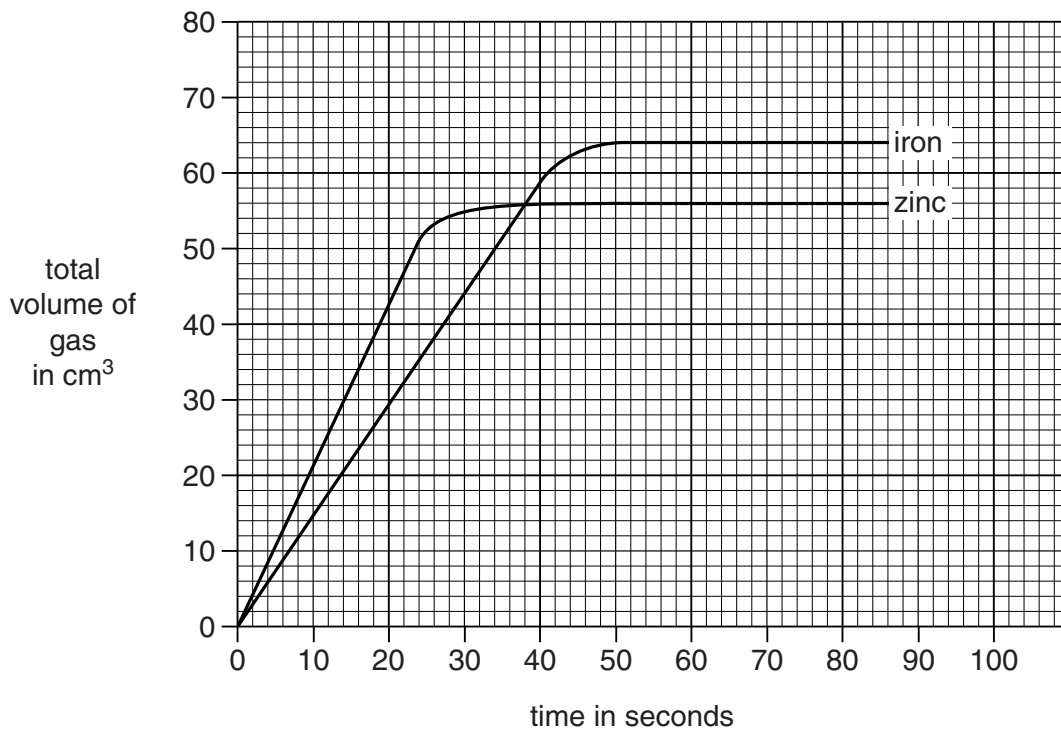
They react dilute hydrochloric acid with zinc powder and then with iron powder.

Look at the apparatus they use.



Every 10 seconds they measure the volume of gas in the gas syringe.

Look at the graph of the results.



- (i) The graph for the reaction of **zinc** is different from the graph for **iron**.

Write about **two** differences in these graphs.

.....

.....

.....

.....

..... [2]

- (ii) Iron powder reacts faster than a lump of iron of the same mass.

Explain why.

.....

..... [1]

- (iii) Fatimah and Sam want to make the reaction between iron powder and dilute hydrochloric acid **faster**.

They do not want to change the mass of the iron powder or the volume of acid.

Write about **three** ways they can make the reaction faster.

.....

.....

.....

.....

.....

.....

..... [3]

[Total: 7]

END OF QUESTION PAPER

ADDITIONAL ANSWER SPACE

If additional space is required, you should use the following lined page(s). The question number(s) must be clearly shown in the margins.

A large area of the page is reserved for writing, featuring a vertical margin line on the left and horizontal dotted lines for writing.

A large rectangular area with a solid vertical line on the left side and horizontal dotted lines extending across the page, providing a grid for writing answers.



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